

LABORATORY CALCULATIONS AND SPECIFIC GRAVITY TABLES. BY JOHN S. ADRIANCE, A.M. Third Edition. Revised and Enlarged. 1897. ix + 144 pp. Price, \$1.25.

This little book, which is well known to many chemists, appears in a new and somewhat enlarged edition.

Many an analyst has felt impatient at the loss of time necessary in making analytical calculations. To such this book will be found useful, containing as it does calculations for most of the reagents employed in analytical chemistry. The percentage and specific gravity tables, which are numerous, as well as the directions for preparing many reagents, will be found convenient.

It is to be regretted that the atomic weights employed by the author could not have been taken from the recent edition of Professor Clarke's "Constants of Nature," instead of his tables published in 1882. The book as it is, however, will be useful to most chemists, and should find a place in every well regulated laboratory.

J. F. MCGREGORY.

AN INTRODUCTORY COURSE OF QUANTITATIVE CHEMICAL ANALYSIS, WITH EXPLANATORY NOTES AND STOICHIOMETRICAL PROBLEMS. BY HENRY P. TALBOT, PH.D., Associate Professor of Analytical Chemistry in the Massachusetts Institute of Technology. 125 pp. New York: The Macmillan Co. Price \$1.50.

There are few readers of analytical chemistry who have not felt the need of a book which shall in a clear and concise style give the principles of quantitative analysis, and sufficient of the practice to enable the student to make his way in the larger and fuller manuals. Such a book is needed not only for those who expect to become analytical chemists, but for the much larger number who pursue chemistry in colleges, merely in course of a general education. It is hardly too much to say that it is difficult to see how a better book could be prepared for this purpose than the one before us. The general directions and the specific directions under each analysis are so clearly given that an instructor could oversee the work of a large number of students, and in each case the theory of every step is so lucidly explained that the book has a distinct educational value. The examples are well chosen and illustrate the most important methods of procedure. They are in gravimetric analysis, chlorine, iron, sulphuric acid, phosphorus pentoxide, calcium and magnesium in dolomite, lead, copper (electrolytic), and zinc in